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Chemistry Stoichiometry Packet Answers

Unit 6 Packet - Page 9 of 12

Stoichiometry - Mass to Mass Problems

Unit 6 Packet - Page 10 of 12

Stoichiometry Worksheet #1 Perform the following calculations. Be sure to use proper units! Answer the following g mol and/or mol g conversion problems. 1.

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How many g in 7.00 mol of N₂? _____ 2.

How many g in 0.455 mol of NaCl? _____

3.

Unit 6 - Stoichiometry Packet

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Answer the following stoichiometry-
related questions: 1) Write the balanced

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equation for the reaction of acetic acid with aluminum hydroxide to form water and aluminum acetate: 2) Using the equation from problem #1, determine the mass of aluminum acetate that can be made if I do this reaction with 125 grams of acetic acid and 275 grams of aluminum hydroxide.

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Stoichiometry Practice Worksheet

W/ answers Website Upload Unit 6:

Reactions and Stoichiometry (Link to
Prentice Hall Text: Chapters 7, 8 & 9)

Name: _____ Date Due Assignments Page

Number: Problem Numbers Assignment

1: Interpreting Equations and Balancing

232: 32, 39

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Unit 6: Reactions and Stoichiometry

Stoichiometry: Problem Sheet 2 Subject:
Chemistry Author: John Bergmann & Jeff
Christopherson Keywords: stoichiometry,
chemical equation Last modified by:
jlbrock Created Date: 4/12/2009 7:15:00
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Chapter 12

These mole ratios are used to solve problems such as how many moles of carbon dioxide, CO_2 , would be produced from 6.25 moles of oxygen gas? Solution: $6.25 \text{ moles O}_2 (3 \text{ mol CO}_2 / 5 \text{ mol O}_2) = 3.75 \text{ moles CO}_2 + +$
YouTube Video: Solving Stoichiometry Problems by weiner7000 STOP at 7:25

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until you have read through the next three sections.

Chapter 13 Stoichiometry

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2018 - this packet is a cumulative review of many topics from the year that are fair game on the stoichiometry test a
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STEP 1 Fold a sheet of paper in half lengthwise.

Chapter 11: Stoichiometry

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PracticePacket((Unit6: Moles(&(Stoichiometry

Stoichiometry is the calculation of quantitative relationships of the reactants and products in chemical

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reactions. Given enough information, we can use stoichiometry to calculate the moles and masses within a chemical equation. In this lesson, we will look into some examples of stoichiometry problems. What a chemical equation tells you?

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