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Acids And Bases

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Circle the ones that are "Soluble". 1. Acid 2. Base 3. Neutral 4.

Neutralize 5. Acid Rain A. To mix acids and bases to cancel each

other out and make salt water. B. A compound that adds H^+ ions to

water. C. Equal number of H^+ and OH^- ions; water is an example.

D.

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Acids and bases Acids are chemicals that add H^+ ions to water

when you make a solution. Bases are chemicals that add OH^- ions to

water in a solution. In other words when you put a chemical into

water and stir it up, if it makes H^+ ions, it is an acid; if it makes

OH^- ions: it is a base. If an acid makes a lot of H^+ ions, we call it a

strong acid, like HCl . If it makes just a few

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Solutions, Acids and Bases : Acids and Bases 1 (Chapter 25)

(Matching). Acids and Bases 2 (Chapter 25) (Matching). Acid,

Base, or Neutral (MC) - learn the differences. Solutions 1 (Chapter

23) (Matching). Solutions 2 (Chapter 23) (Matching). Unsaturated,

Saturated, or Supersaturated (Multiple Choice). Solubility Graphs -

Although this is covered more in Chemistry than IPC, it has been on

TAKS ...

Murray IPC homework

HW5:4 Acids and Bases Mr. Murray, IPC cstephenmurray.com 1.

What is an acid? 2. What is a base? 3. Are foods acidic or basic? 4.

Are cleaning products acidic or basic? 5. Why? 6. What is pH? 7. If

you needed a pH of 7.8 in a pool, but the pool has a pH of 6.8, what

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do you add to the pool? 8. Acid, Base, or Neutral? 9. What is a titration? 10.

HW5:4 Acids and Bases Assigned: Thurs., 12/7 and Fri., 12 ...

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Bases are compounds that add OH^- ions to water when in a solution. Cl^- H^+ O^- H^+ OH^- ion NaOH — sodium hydroxide: a very strong base. O^- H^+ Na^+ Na^+ In water it breaks up (dissociates) adding OH^- ions to the water. Cl^- H^+ In water HCl — Hydrochloric acid: a very strong acid. In water it breaks up (dissociates) and adds H^+ ions.

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Where To Download Stephen Murray Acids And Bases Answer Keythat adds a lot of H^+ ions to water. pH and Acid Rain Plants and animals need water Acids and Bases - Mr Murray's Science and Music Acids and bases Acids are chemicals that add H^+ ions to water when you make a solution. Bases are chemicals that add OH^- ions to water in a solution.

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Naming Compounds

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Acids are substances which produce hydrogen ions in solution. Bases are substances which produce hydroxide ions in solution. Neutralisation happens because hydrogen ions and hydroxide ions react to produce water. Hydrochloric acid is neutralised by both sodium hydroxide solution and ammonia solution.

THEORIES OF ACIDS AND BASES - chemguide

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Hydrosulfuric Acid- Nitric Acid- Sulfuric Acid- Acetic Acid- Boric Acid- Give the formula for the following bases: Sodium Hydroxide- Potassium Hydroxide- Calcium Hydroxide- Iron (II) Hydroxide- 4. <http://chemistry.about.com/od/acidsbases/a/acidsbaseterms.htm> Scroll down to Properties of Acids. Complete the following sentences for Acids

Introduction to Acids & Bases: A WebQuest

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Chapter 22 - 25 Review Vocabulary - Know the following vocabulary and how to apply them. 1. Neutralize 2. Neutral 3. Acid Rain 4. Acid 5. Base A. When pollution causes rain to be acidic (pH of less than 5.6). B. To mix acids and bases to cancel each other out and make salt water.

Chapter 22 - 25 Review - Mr Murray's Science and Music

between an acid and a base. E. A compound that adds a few H^+ ions to water. F. A compound that adds a lot of H^+ ions to water. pH and Acid Rain Plants and animals need water close to neutral (pH 7) to survive. Due to pollution from combustion reactions, rain today can be acidic. Rain less than pH 5.6 we call acid rain. Acid rain can kill plants, cause

Strong acids

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