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*Solubility Product Constant
(K_{sp})* ~~Solubility Problems
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$x = 1.33 \times 10^{-5} \text{ M}$. This is the answer because there is a one-to-one relationship between the Ag⁺ dissolved and the AgCl it came from.

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So, the molar solubility of AgCl is 1.33×10^{-5} moles per liter. Calculate the molar solubility (in mol/L) of a saturated solution of the substance.

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Answer. First, treat the solubility equilibrium in units of molarity and then change the concentration to ppb.

$$\text{Cu}_3(\text{AsO}_4)_2(s) \rightarrow 3\text{Cu}^{2+}(\text{aq}) + 2\text{AsO}_4^{3-}(\text{aq})$$

$$K_{sp} = [\text{Cu}^{2+}]^3 [\text{AsO}_4^{3-}]^2 = 7.6 \times 10^{-36}$$

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Initial 0 0. Change +3x +2x.
Equilibrium 3x 2x. $7.6 \times 10^{-36} = [3x]^3 [2x]^2 = 108x^5$
 $x = 3.7 \times 10^{-8} \text{ M} = \text{molar solubility of copper(II) arsenate.}$

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Answers. 1 Solubility Curves
. There are charts and
tables available that we can
use to get an idea of how
soluble a certain solute is
in a certain solvent. We
will take a look at two of

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them in these next two sections. Solubility curves, like the one shown here, tell us what mass of solute will dissolve in 100g (or 100mL; see note

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Key ~~...~~

very small (the solubility is reduced in the presence of a common ion), the term " $0.020 + x$ " is the same as " $0.020.$ " (You can leave x in the term and use the quadratic equation but it

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will not improve the significance of your answer.) : $1.1 \times 10^{-10} = [x][0.020 + x] = [x][0.020]$
 $x = 5.5 \times 10^{-9} \text{ M}$ Effect of the Common Ion on Solubility

~~Unit 12 Subjects SOLUBILITY~~

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~~PRODUCT CALCULATIONS~~

now take the solubility, and multiply it by 325 and divide it by 100 (Rule of three). A few tips for the rest of the problem: 2) Procedure is in reverse order to 1) 3) Subtract the

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solubilities of...

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Sample Problem #2 If 0.0067g CaCO₃ soluble in 1.0L of water, calculate K_{sp} molar solubility = (0.0067g/L)(1

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$\text{mol}/100\text{g}) = 6.7 \times 10^{-5} \text{ M CaCO}_3(\text{s})$
 $\text{Ca}^{+2} + \text{CO}_3^{-2}$
 $6.7 \times 10^{-5} \text{M} \quad 6.7 \times 10^{-5} \text{M}$
 $K_{\text{sp}} = [\text{Ca}^{+2}][\text{CO}_3^{-2}] =$
 $[6.71 \times 10^{-5}][6.7 \times 10^{-5}] =$
 4.5×10^{-9} Sample Problem #3
 If 0.017g CaF_2 soluble in 1.0L of water, calculate K_{sp}

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~~Ksp Problems — Chemistry~~
SOLUBILITY CURVES Answer the following questions based on the solubility curve below. Which salt is least soluble in water .. 2. How many grams of potassium chloride

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can be dissolved in 200 g of water at 80° C? 10 3. At 40° C, how much potassium nitrate can be dissolved in 300 g of water? 4. Which salt shows the least change

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Solubility Curve Practice
Problems Answer Key
SOLUBILITY PROBLEMS. Here
are some practice problems
for writing K_{sp}

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expressions. Write the chemical equation showing how the substance dissociates and write the K_{sp} expression: PART 1: 1) $AlPO_4$ 4 2) $BaSO_4$ 4 3) CdS 4) $Cu_3(PO_4)_2$ 5) $CuSCN$ 6) Hg_2Br_2 7) $AgCN$ 8) $Zn_3(AsO_4)_2$

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4) 2 9) Mn(IO 3) 2 10 ...

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\therefore Molar solubility =
 7.07×10^{-7} . 5. Ag₂CO₃
 $\rightleftharpoons 2\text{Ag}^+ + \text{CO}_3^{2-}$ - $K_{sp} =$

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$8.1 \times 10^{-12} = 4x^3$. $2x \cdot x \cdot K_{sp} = [Ag^+]^2 [CO_3^{2-}]$
 $8.1 \times 10^{-12} = 4x^3$
 $x^3 = 2.025 \times 10^{-12}$. $x = 1.3 \times 10^{-4}$ \therefore Molar solubility is 1.3×10^{-4} .
 6. $AgI \rightleftharpoons Ag^+ + I^-$
 $NaI \Rightarrow Na^+ + I^-$
 $0.2 \quad 0.2$

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Solubility Graph Worksheet
Answers Solubility Product
Worksheet - Answers. 1) What
is the concentration of a
saturated silver (I) acetate
solution? $K_{sp}(\text{AgC}_2\text{H}_3\text{O}_2)$

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2) = 1.94×10^{-3} . Since K_{sp}
= $[Ag^+][C_2H_3O_2^-]$, and
the concentration of silver
ions is the same as the
concentration of acetate
ions, we can set up the

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